

Calculation practice for PAL groups

How many moles of NaCl are present in 500ml of a 0.1M solution of this salt?

How many grams of ammonium sulphate are required to make 1L of a 0.5M solution (mw ammonium sulphate = 132.14)?

How many moles are present in 36g of NaCl (mw = 58.44)

Find the number of mmoles of EDTA in 10 μ l of a 0.2M solution.

You want to make 0.2L of a 0.3% w/v MgCl_2 solution. How many

(a) grams and

(b) moles of this salt will this contain?

What is the pH of 0.01M HCl?

What is the pH of 0.45M HCl?

What is the pH of 0.05M KOH?

What is the pH of 0.055M NaOH?

What is the proton concentration of a solution of pH 2.83? State your units.